Stephen Heuer, Nutripath - Synergistic Nutrition <stephen@sygn44.com>

Monoatomic Minerals

1 message

Ray <raybse@cableone.net> Mon, Jun 17, 2024 at 12:15 AM
Reply-To: raybse@cableone.net
To: Stephen Heuer <stephen@sygn44.com>

Dear Stephen,

Monoatomic Elements

Monoatomic elements are nothing more than elements which are chemically isolated, i.e. instead of 60 atoms of Carbon are 34 atoms of Silicon being bound together in something called a Buckministerfullerene or a knobbier version of the same. The significance lies in the fact that when a single element metal progresses from a normal metallic state to a monoatomic state, it passes through a series of chemically different states. These include:

· An alloy of numerous atoms of the same element, which exhibit all the characteristics normally associated with the metal: electrical conductivity, color, specific gravity, density, and so forth. The atom's intrinsic temperature might be room temperature.

· A combination of significantly fewer atoms of the same element, which no longer exhibit all of the characteristics normally associated with the metal. For example, the electrical conductivity or color might change. The atom's intrinsic temperature drops, for example, to 50 to 100 oK (or about two hundred degrees below zero oC).

· A Microcluster of far few atoms -- typically on the order of less than one hundred atoms, and as few as a dozen or so atoms. The metal characteristics begin to fall off one by one until the so-called metal is hardly recognized. The intrinsic temperature has now fallen to the range of 10 to 20 oK, only slightly above Absolute Zero.

· A Monoatomic form of the element -- in which each single atom is chemically inert and no longer possesses normal metallic characteristics; and in fact, may exhibit extraordinary properties. The atom's intrinsic temperature is now about 1 oK, or close enough to Absolute Zero that Superconductivity is a virtually automatic condition.

A case in point is Gold. Normally a yellow metal with a precise electrical conductivity and other metallic characteristics, the metallic nature of gold begins to change as the individual gold atoms form chemical combinations of increasingly small numbers. At a microcluster stage, there might be 13 atoms of gold in a single combination. Then, dramatically, at the monoatomic state, gold becomes a forest green color, with a distinctly different chemistry. It's electrical conductivity goes to zero even as its potential for Superconductivity becomes maximized. Monoatomic gold can exhibit substantial variations in weight, as if it were no longer fully extant in space-time.

Other elements which have many of these same properties are the Precious Metals, which include Ruthenium, Rhodium, Palladium, Silver, Osmium, Iridium, Platinum, and Gold. All of these elements have to greater or lesser degree, the same progression as gold does in continuously reducing the number of atoms chemically connected. Many of these precious elements are found in the same ore deposits, and in their monoatomic form are often referred to as the White Powder of Gold.

Monoatomic elements apparently exist in nature in abundance. Precious Metal ores are, however, not always assayed so as to identify them as such. Gold miners, for example, have found what they termed “ghost gold” -- “stuff” that has the same chemistries as gold, but which were not yellow, did not exhibit normal electrical conductivity, and were not identifiable with ordinary emission spectroscopy. Thus they were more trouble than they were worth, and generally discounted.

However, in a technique called “fractional vaporization”, the monoatomic elements can be found and clearly identified via a more advanced emission spectroscopy. This fact was first discussed by David Radius Hudson, who was attempting to separate gold and silver from raw ore -- but was hindered by the ghost gold which had no apparent intrinsic value.

The process involved placing a sample on a standard carbon electrode, running a second carbon electrode down to a position just above the first, and then striking a Direct Current arc across the electrodes. The electrical intensity of the arc would ionize the elements in the sample such that each of the elements would give off specific, identifying frequencies of light. By measuring the specific frequencies of light (the spectrum of the element or elements), one could then identify which elements were in the sample. Typically, such spectroscopic analysis involves striking the arc for 10 to 15 seconds, at the end of which, the carbon electrodes are effectively burned away. According to the majority of American spectroscopists, any sample can be ionized and read within those 15 seconds.

In the advanced technique, the carbon electrodes are sheathed with an inert gas (such as Argon). This allows the emission spectroscopy process to be continued far beyond the typical 15 seconds, in order to fully identify all of the elements in their various forms.

When this was done, in the first seconds, the ghost gold might be identified as iron, silicon, and aluminum. But as the process continued for as long as 300 seconds, palladium began to be read at about 90 seconds, platinum at 110 seconds, ruthenium at 130 seconds, rhodium at 145 seconds, iridium at 190 seconds, and osmium at 220 seconds. These latter readings were the monoatomic elements. Commercially available grades of these metals were found to be including only about 15% of the emission spectroscopic readings.

The mining activity of what is considered the best deposit in the world for six of these elements (Pd, Pt, Os, Ru, Ir, and Rh) yields one-third of one ounce of all these precious metals per ton of ore. But this is based on the standard spectroscopic analysis. When the burn is continued for up to 300 seconds, the same ores might easily yield emission lines suggesting: 6 to 8 ounces of palladium, 12 to 13 ounces of platinum, 150 ounces of osmium, 250 ounces of ruthenium, 600 ounces of iridium, and 1200 ounces of rhodium! Over 2200 ounces per ton, instead 1/3 of 1 ounce per ton! [Keep in mind that rhodium typically sells for $3,000/ounce, while gold sells for $300/ounce!]

The distinguishing characteristic between the first and second readings of the emission spectroscopy for the precious metals is that all of them come in two basic forms. The first is the traditional form of metals: yellow Gold, for example. The second is the very non-traditional form of the metal: the monoatomic state. The chemistries and physics of these two different states of these metals are radically different. More importantly, when the atoms are in the monoatomic state, things really begin to get interesting!

A key to understanding monoatomic elements is to recognize that the monoatomic state results in a rearrangement of the electronic and nuclear orbits within the atom itself. This is the derivation of the term: Orbitally-Rearranged Monoatomic Element (ORME).

A monoatomic state implies a situation where an atom is “free from the influence of other atoms.” Is this, perhaps, a violation of some very basic, absolutely fundamental law of the universe -- which says that nothing is separate? If such a law constituted reality, then a necessary condition for monoatomic elements to even exist would require them to be superconductive, just in order to link them through all distance and time to other superconducting monoatomic elements. This would be necessary in order to prevent separation. The question is whether separation is but the Ultimate Illusion?

Ray Hamilton
P. O. Box 513 Caldwell ID 83606

Blessings Ray Hamilton
www.eBibleProductions.com
www.OrmusMinerals.com
OceanNectar.com
OrmusBrain.com
OrmusForPlants.com
OrmusProbiotics.com
QuantumEnergyOrmus.com
ILoveOrmus.com
LifeForceMinerals.com
MagnesiumBenefits.biz
MagnesiumDew.com
OrmusGoldPowder.com
OrmusWhiteGold.com
WhatIsOrmus.com